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1. $5.3 \times 10^3 \text{ C}$ 3. (a) reduction; (b) oxidation; (c) oxidation; (d) reduction 5. (a) $\text{F}_2 + \text{Ca} \rightarrow \text{Ca}^{2+} + 2\text{F}^-$; $\text{F}_2 + \text{Ca} \rightarrow \text{Ca}^{2+} + 2\text{F}^-$; (b) Considerations include: cost of the materials used in the battery, toxicity of the various components (what constitutes proper disposal), should it be a primary or secondary battery, energy requirements (the "size" of the battery/how long ...

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reversible reaction. law of chemical equilibrium. equilibrium constant expression. $K_{eq} > 1$. chemical reaction that can occur in both the forward and the reverse. ... particular ratio of reactant and product concentrations has a ... $keq = \frac{[C]^c [D]^d}{[A]^a [B]^b}$.

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17-7 K and the extent of reaction K reflects a particular ratio of product concentrations to reactant concentrations for a reaction. A small value for K indicates that the reaction yields little product before reaching equilibrium. The reaction favors the reactants. K therefore indicates the extent of a reaction, i.e., how far a reaction proceeds towards the products at a given

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17.6 Corrosion – Chemistry
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Electrolysis. Another chapter in this text introduced the chemistry of reduction-oxidation (redox) reactions.

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17-1 CHAPTER 17 EQUILIBRIUM: THE EXTENT OF

CHEMICAL REACTIONS 17.1 If the rate of the forward reaction

exceeds the rate of reverse reaction, products are formed faster than they are consumed. The change in reaction conditions results in more products and less reactants. A change in reaction

CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS

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Organic chemistry is defined as the study of compounds of carbon or the chemistry of hydrocabons and its derivatives. The term

‘ organic ’ is misleading. Earlier the term organic chemistry was used

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to describe the study of compounds obtained from living organisms, while the term inorganic chemistry was used for the study of compounds obtained ...

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Chemistry, 11e (Brown/LeMay/Brusten/Murphy) Chapter 17: Additional Aspects of Aqueous Equilibria 8) Calculate the pH of a solution prepared by dissolving 0.250 mol of benzoic acid $\text{C}_6\text{H}_5\text{COOH}$ and 0.150 mol of sodium benzoate $\text{C}_6\text{H}_5\text{COO}^- \text{Na}^+$ in water sufficient to yield 1.00 L of solution. The K_a of benzoic acid is 6.50×10^{-5} .

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