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Acids and Bases

Section 118.18.1

Figure 18.1

Rhododendrons

flourish in rich, moist
soil that is moderately
acidic, whereas

semper-vivum,

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commonly called hen
and chicks, grows
best in drier, slightly
basic soil.

Introduction to Acids
and

Bases-!).)DEADifferent
models help describe
the behavior of acids
and bases.

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AP Chemistry—
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CHAPTER 16 STUDY GUIDE- Acid-Base Equilibrium 16.1

Acids and Bases: A
Brief Review • Acids
taste sour and cause
certain dyes to
change color. • Bases
taste bitter and feel
soapy. • Arrhenius
concept of acids and
bases: • An acid is a
substance that, when
dissolved in water,

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increases the
concentration of H^+
ions.

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18.2 The strong acid
hydrochloric acid
(HCl), also called
muriatic acid,

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Figure 18.1

Rhododendrons

flourish in

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acid–base condition.
In their classic article,
Pe-ters and Van Slyke
defined acid–base
balance in the blood

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as the chemical state resulting from the balance between cations and anions.¹² Carrying this idea to the extreme, one could view metabolic acid–base disorders as the predicted consequences of primary fluid and electrolyte imbalance.

Integration of Acid

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Base and Electrolyte Disorders

Chapter 18 - Water, Electrolyte, and Acid-Base Balance To be in balance, the quantities of fluids and electrolytes leaving the body should be equal to the amounts taken in. Electrolytes are molecules that release ions in water.

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Anything that alters the concentrations of electrolytes will also alter the concentration of water, and vice versa.

Chapter Eighteen
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Chapter 18 Water ...
segment 18 Chap 4:
Electrolytes and Acid-
Base Reactions blem
4.37 Classify each of

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the following

substances as a nonelectrolyte, weak electrolyte, or strong electrolyte in water.

NH_4OH , HF , KClO_4 ,
 $\text{Cu}(\text{NO}_3)_2$

nonelectrolyte weak
electrolyte strong
electrolyte blem 4.38

- Enhanced - with
Feedback Classify
each of the following
aqueous solutions as

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a nonelectrolyte,
weak electrolyte, or
strong electrolyte.

Solved: Signment 18

Chap 4: Electrolytes

And Acid-Base Rea ...

Reaction of Acids and
Bases with each other

- Acids and Bases
react to form salt and
water. Acid + Base

Salt + H₂O •

Neutralisation

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Reaction: Reaction of acid with a base is called as neutralization

reaction. Example:



H_2O • Strong Acid + Weak Base

Acidic salt + H_2O •

Weak Acid + Strong

Base Basic salt + H_2O ...

Notes of Ch 2 Acids,

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Bases and Salts |

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Mineral Acids - Acids that prepare from mineral are known as mineral acids/
inorganic acids/man-made acids or
synthetic acid, such as hydrochloric acid (HCl), sulphuric acid (H₂SO₄), nitric acid (HNO₃), etc. Note: Organic acids are

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always weak but minerals acids can be strong as well as weak. We can take dilute weak acids (like organic acid) in our body.

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Chemistry Companion Course helps students learn the essential lessons associated with acids, bases and salts.

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Chemistry Chapter 19: Acids, Bases and Salts ...

Chapter 15 – Acid-Base Equilibria. Acid-Base Equilibria . 15.1 Solutions of Acids or Bases Containing a Common Ion . A. Common Ion 1. Ion provided in solution by an aqueous acid (or base) as well as a salt a. $\text{HF}(\text{aq})$ and NaF

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(F-in common) HF(aq)

$\text{H}^+(\text{aq}) + \text{F}^-(\text{aq})$

Excess F added by

NaF. Equilibrium

shifts away from

added . component.

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